Heat of Formation Quiz

Correct answer gets full marks. Anything else, more work shown = more part marks.

Many cigarette lighters contain liquid butane. C4H10 (l). Using enthalpies of formation (from your handout) calculate the quantity of heat produced when 46.0g of butane is completely burned in air.

(∆H˚f – Butane(l) = –147.6kJ/mol)?

The quantity of heat produced is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_